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9204/CC

Paper 2 Chemistry Core

Mark scheme

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Mark schemes are prepared by the Lead Assessment Writer and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation events which all associates participate in and is the scheme which was used by them in this examination. The standardisation process ensures that the mark scheme covers the students' responses to questions and that every associate understands and applies it in the same correct way. As preparation for standardisation each associate analyses a number of students' scripts. Alternative answers not already covered by the mark scheme are discussed and legislated for. If, after the standardisation process, associates encounter unusual answers which have not been raised they are required to refer these to the Lead Examiner.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of students' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

Further copies of this mark scheme are available from oxfordaqaexams.org.uk

Level of response marking instructions

Level of response mark schemes are broken down into levels, each of which has a descriptor. The descriptor for the level shows the average performance for the level. There are marks in each level.

Before you apply the mark scheme to a student's answer read through the answer and annotate it (as instructed) to show the qualities that are being looked for. You can then apply the mark scheme.

Step 1 Determine a level

Start at the lowest level of the mark scheme and use it as a ladder to see whether the answer meets the descriptor for that level. The descriptor for the level indicates the different qualities that might be seen in the student's answer for that level. If it meets the lowest level then go to the next one and decide if it meets this level, and so on, until you have a match between the level descriptor and the answer. With practice and familiarity you will find that for better answers you will be able to quickly skip through the lower levels of the mark scheme.

When assigning a level you should look at the overall quality of the answer and not look to pick holes in small and specific parts of the answer where the student has not performed quite as well as the rest. If the answer covers different aspects of different levels of the mark scheme you should use a best fit approach for defining the level and then use the variability of the response to help decide the mark within the level, ie if the response is predominantly level 3 with a small amount of level 4 material it would be placed in level 3 but be awarded a mark near the top of the level because of the level 4 content.

Step 2 Determine a mark

Once you have assigned a level you need to decide on the mark. The descriptors on how to allocate marks can help with this. The exemplar materials used during standardisation will help. There will be an answer in the standardising materials which will correspond with each level of the mark scheme. This answer will have been awarded a mark by the Lead Examiner. You can compare the student's answer with the example to determine if it is the same standard, better or worse than the example. You can then use this to allocate a mark for the answer based on the Lead Examiner's mark on the example.

You may well need to read back through the answer as you apply the mark scheme to clarify points and assure yourself that the level and the mark are appropriate.

Indicative content in the mark scheme is provided as a guide for examiners. It is not intended to be exhaustive and you must credit other valid points. Students do not have to cover all of the points mentioned in the Indicative content to reach the highest level of the mark scheme.

An answer which contains nothing of relevance to the question must be awarded no marks.

Information to Examiners

1. General

The mark scheme for each question shows:

- the marks available for each part of the question
- the total marks available for the question
- the typical answer or answers which are expected
- extra information to help the Examiner make his or her judgement
- the Assessment Objectives, level of demand and specification content that each question is intended to cover.

The extra information is aligned to the appropriate answer in the left-hand part of the mark scheme and should only be applied to that item in the mark scheme.

At the beginning of a part of a question a reminder may be given, for example: where consequential marking needs to be considered in a calculation; or the answer may be on the diagram or at a different place on the script.

In general the right-hand side of the mark scheme is there to provide those extra details which confuse the main part of the mark scheme yet may be helpful in ensuring that marking is straightforward and consistent.

2. Emboldening and underlining

2.1 In a list of acceptable answers where more than one mark is available 'any **two** from' is used, with the number of marks emboldened. Each of the following bullet points is a potential mark.

2.2 A bold **and** is used to indicate that both parts of the answer are required to award the mark.

2.3 Alternative answers acceptable for a mark are indicated by the use of **or**. Different terms in the mark scheme are shown by a / ; eg allow smooth/free movement.

2.4 Any wording that is underlined is essential for the marking point to be awarded.

3. Marking points

3.1 Marking of lists

This applies to questions requiring a set number of responses, but for which students have provided extra responses. The general principle to be followed in such a situation is that 'right + wrong = wrong'.

Each error/contradiction negates each correct response. So, if the number of errors/ contradictions equals or exceeds the number of marks available for the question, no marks can be awarded.

However, responses considered to be neutral (indicated as * in example 1) are not penalised.

Example 1: What is the pH of an acidic solution?

[1 mark]

Student	Response	Marks awarded
1	green, 5	0
2	red*, 5	1
3	red*, 8	0

Example 2: Name two planets in the solar system.

[2 marks]

Student	Response	Marks awarded
1	Neptune, Mars, Moon	1
2	Neptune, Sun, Mars, Moon	0

3.2 Use of chemical symbols/formulae

If a student writes a chemical symbol/formula instead of a required chemical name, full credit can be given if the symbol/formula is correct and if, in the context of the question, such action is appropriate.

3.3 Marking procedure for calculations

Marks should be awarded for each stage of the calculation completed correctly, as students are instructed to show their working. Full marks can, however, be given for a correct numerical answer, without any working shown.

3.4 Interpretation of 'it'

Answers using the word 'it' should be given credit only if it is clear that the 'it' refers to the correct subject.

3.5 Errors carried forward

Any error in the answers to a structured question should be penalised once only.

Papers should be constructed in such a way that the number of times errors can be carried forward is kept to a minimum. Allowances for errors carried forward are most likely to be restricted to calculation questions and should be shown by the abbreviation ecf in the marking scheme.

3.6 Phonetic spelling

The phonetic spelling of correct scientific terminology should be credited **unless** there is a possible confusion with another technical term.

3.7 Brackets

(…..) are used to indicate information which is not essential for the mark to be awarded but is included to help the examiner identify the sense of the answer required.

3.8 Allow

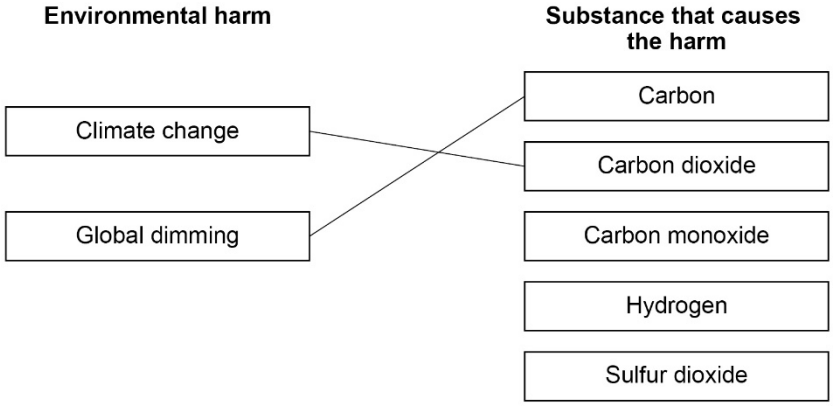
In the mark scheme additional information, 'allow' is used to indicate creditworthy alternative answers.

3.9 Ignore

Ignore is used when the information given is irrelevant to the question or not enough to gain the marking point. Any further correct amplification could gain the marking point.

3.10 Do <u>not</u> accept

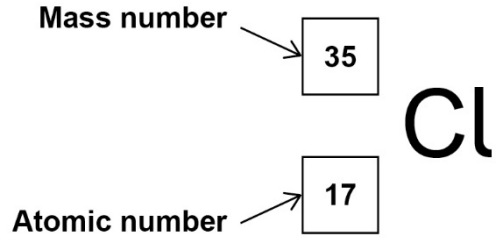
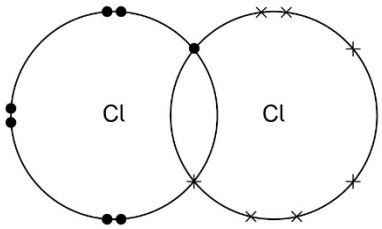
Do **not** accept means that this is a wrong answer which, even if the correct answer is given as well, will still mean that the mark is not awarded.

Question	Answers	Extra information	Mark	AO/Spec. Ref.
01.1	mixture		1	AO1 3.16.1.1a
01.2	fractional distillation		1	AO1 3.16.1.1c
01.3	C_3H_8		1	AO1 3.16.1.2b
01.4	 do **not** accept more than one line from a box on the left		1 1	AO1 × 2 3.16.1.2d
01.5	as relative formula mass increases viscosity increases B does not fit the trend	allow M_r for relative formula mass allow B has a high viscosity and a low relative formula mass allow viscosity of B is higher than expected allow relative formula mass of B is lower than expected	1 1	AO3 × 2 3.16.1.2c
01.6	decreases		1	AO1 3.16.1.2c
Total			**8**	

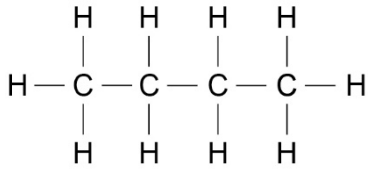
Question	Answers	Extra information	Mark	AO/Spec. Ref.
02.1	4(%)		1	AO2 3.11
02.2	(mass =) $\frac{300 \times 6}{100}$ = 18 (kg)		1 1	AO2 AO3 3.11
02.3	hydrogen ion		1	AO1 3.11.1f
02.4	**Solution**: Nitric acid; Alkaline solution **Equipment used to measure volume of solution**: Beaker; Burette; Funnel; Pipette; Ruler do **not** accept more than one line from a box on the left		1 1	AO1 × 2 3.12.4b
02.5	neutralised		1	AO1 3.11.1b 3.12.4b
02.6	(mean =) $\frac{24.35 + 24.25 + 24.35 + 24.30}{4}$ = 24.31 (cm^3)	 allow 24.312 / 24.3125 (cm^3)	1 1	AO2 AO3 3.12.4b
02.7	ammonia	allow ammonium hydroxide	1	AO2 3.11.1d
02.8	9		1	AO2 3.11.1d
Total			**11**	

Question	Answers	Extra information	Mark	AO/Spec. Ref.
03.1	(volume =) $\frac{0.015}{0.15}$		1	AO2 × 3 3.12.4a
	= 0.1 (dm^3)		1	
	= 100 (cm^3)	allow correct conversion of an incorrectly calculated volume	1	
03.2	silver iodide	allow AgI	1	AO2 3.12.1a
03.3	filtration	allow filter(ing)	1	AO4 3.10.1c
03.4	lilac		1	AO3 3.10.2a
03.5	warm / heat the water	allow stir the mixture allow shake the tube	1	AO4 3.11.2b
03.6	sodium hydroxide solution		1	AO3 3.10.2c
03.7	bromide		1	AO3 3.10.2e
Total			**9**	

Question	Answers	Extra information	Mark	AO/Spec. Ref.
04.1	electrons		1	AO2 3.9.2e
04.2	hydrogen is less reactive than sodium		1	AO2 3.9.2g
04.3	(chloride ions are) negatively charged (so) are attracted (to the positive electrode)		1 1	AO2 × 2 3.9.2c
04.4	**Substance** / **Product made from the substance** Chlorine → Plastic Sodium hydroxide → Soap do **not** accept more than one line from a box on the left		1 1	AO1 × 2 3.9.2j
04.5	universal indicator (colour change from) green (to) blue / purple	MP2 and MP3 must be consistent with indicator given	1 1 1	AO1 AO3 × 2 3.11.1f
Total			**9**	

Question	Answers	Extra information	Mark	AO/Spec. Ref.
05.1	liquid		1	AO3 3.7.1a
05.2	(from) orange (to) colourless	 ignore clear	1 1	AO1 × 2 3.16.1.3d
05.3			1 1	AO2 × 2 3.7.2fh
05.4	covalent		1	AO1 3.8.1g
05.5	 1 pair of bonded electrons 6 non-bonded electrons on both Cl atoms	allow any combination of circles, dots, crosses or $e^{(-)}$ an answer of  scores **2** marks	 1 1	AO2 × 2 3.8.1g
05.6	iodine potassium bromide		1 1	AO2 AO3 3.8.1e 3.13.1e
Total			**10**	

Question	Answers	Extra information	Mark	AO/Spec. Ref.
06.1	zinc chloride		1	AO2 3.11.2a
06.2	hydrogen		1	AO1 3.9.1.1a
06.3	(mass =) 0.0040 × 65 = 0.26 (g)		1 1	AO2 × 2 3.12.3a
06.4	(rate =) $\frac{48}{120}$ = 0.4 (cm^3/s)		1 1	AO2 AO3 3.14.1a
06.5	(most reactive) magnesium (zinc) iron		1	AO3 3.9.1.1a
06.6	copper does not react (with dilute hydrochloric acid)		1	AO3 3.9.1.1a
06.7	(bronze and gunmetal are) mixtures of metals		1 1	AO1 AO3 3.9.1b
06.8	increases strength		1	AO3 3.9.1b
06.9	(material) poly(ethene) (reason) does not react with (hydrochloric) acid	allow copper	1 1	AO3 × 2 3.16.2d
Total			**13**	

Question	Answers	Extra information	Mark	AO/Spec. Ref.
07.1	C_5H_{12}		1	AO2 3.16.1.2a
07.2			1	AO2 3.8.3a 3.16.1.2b
07.3	C_2H_4		1	AO2 3.16.1.3b
07.4	one bond in each molecule of **Z** is a double bond		1	AO1 3.16.1.3b
07.5	vaporises		1	AO4 3.16.1.3a
07.6	$C_{10}H_{22}$ is a liquid	allow $C_{10}H_{22}$ is not a gas	1	AO3 3.7.1a
07.7	(use for $C_{10}H_{22}$) fuel (use for **Z**) making poly(ethene)	 allow making polymer / plastic	1 1	AO2 3.16.1.3e 3.16.2a
Total			**8**	

Question	Answers	Extra information	Mark	AO/Spec. Ref.
08.1	rubidium		1	AO1 3.13.1a
08.2	chlorine		1	AO2 3.7.2i
08.3	helium		1	AO1 3.7.3c
08.4	potassium	allow K	1	AO2 3.8.1c
08.5	36		1	AO2 3.8.1e
08.6	(their atoms) have stable arrangements of electrons	allow (their atoms) have full outer energy levels allow (their atoms) have full outer shells	1	AO1 3.7.3c
08.7	(conversion 1.50 kg =) 1500 g (solubility =) $43.2 \times \frac{100}{1500}$ = 2.88 (mol/100 g)	 allow correct use of incorrectly / not converted values from step 1	1 1 1	AO2 × 3 3.8.1d
Total			**9**	

Question	Answers	Extra information	Mark	AO/Spec. Ref.
09.1	$4NH_3$ + **5** O_2 → 4NO + $6H_2O$	allow multiples	1	AO2 3.12.1a
09.2	gain of oxygen by ammonia / nitrogen / hydrogen	allow **1** mark for gain of oxygen	2	AO1 AO2 3.9.1.1d
09.3	activation energy		1	AO1 3.15.2a
09.4	particles have more energy (so) when the particles collide there is enough energy to react **or** (so) when the particles collide the energy is greater than / equal to the activation energy **alternative approach:** more particles have energy greater than / equal to the activation energy (1) (so) particles react when they collide (1)		1 1	AO1 × 2 3.14.1b
09.5	(level of) products is below (level of) reactants	allow the energy decreases (overall) allow energy is transferred to the surroundings ignore references to bond making / breaking	1	AO3 3.15.2a
09.6	(a catalyst) increases the rate of the reaction (by) providing a different reaction pathway **or** (by) lowering the activation energy	allow (a catalyst) speeds up the reaction	1 1	AO1 × 2 3.14.1h 3.15.2e

09.7	low boiling point non-conductor of electricity		1 1	AO1 × 2 3.8.2ce
09.8	(M_r = 14 + (16 × 2) =) 46 (% by mass =) $\left(\frac{14}{46}\right)$ × 100 = 30.43 = 30.4 (%)	 allow correct use of an incorrectly calculated value of M_r allow 30.43478261 allow a correctly calculated answer given to 3 significant figures from an incorrect attempt at the percentage by mass of nitrogen	1 1 1 1	AO2 × 4 3.12.2b

Total			**15**

Question	Answers	Extra information	Mark	AO/Spec. Ref.
10.1	$CaCO_3$ (s)		1 1	AO2 × 2 3.11.1e 3.12.1a 3.12.1b

Question	Answers	Mark	AO/Spec. Ref.
10.2	**Level 3:** The design/plan would lead to the production of a valid outcome. All key steps are identified and logically sequenced.	5–6	AO4 × 6 3.9.1.2a
	Level 2: The design/plan would not necessarily lead to a valid outcome. Most steps are identified, but the plan is not fully logically sequenced.	3–4	
	Level 1: The design/plan would not lead to a valid outcome. Some relevant steps are identified, but links are not made clear.	1–2	
	No relevant content	0	
	Indicative content • add known mass of (Group 2) carbonate to boiling tube • heat the (Group 2) carbonate • measure time for limewater to change using stop clock **or** use of alternative methods such as counting bubbles or comparison of loss of mass of boiling tube and contents • repeat using other (Group 2) carbonates • same mass of each (Group 2) carbonate • same sized pieces of each (Group 2) carbonate • fresh limewater for each experiment • same volume of limewater • judge same degree of cloudiness of limewater such as time for 'x' to disappear • repeat each experiment • the greater the time the slower the rate of reaction		

Total			8