

OCR (A) Chemistry GCSE

Topic 1 - Particles

Flashcards



What are the three states of matter?



What are the three states of matter?

Solid

Liquid

Gas



Describe the arrangement and movement of particles in solids



Describe the arrangement and movement of particles in solids

The particles are packed closely together in a regular arrangement.

The particles vibrate in fixed positions.



Describe the arrangement and movement of particles in liquids



Describe the arrangement and movement of particles in liquids

The particles are close together but able to move past each other.

The particles vibrate and move around each other.



Describe the arrangement and movement of particles in gases



Describe the arrangement and movement of particles in gases

The particles are well separated with no regular arrangement.

The particles vibrate and move freely at high speeds.



How do the relative energies of particles in a solid, liquid and gas compare?



How do the relative energies of particles in a solid, liquid and gas compare?

Particles in a solid have the least amount of energy.

Particles in a gas have the most energy.



What are the names for the state changes from solid to liquid and vice versa?



What are the names for the state changes from solid to liquid and vice versa?

Solid → liquid: Melting

Liquid → solid: Freezing



What are the names for the state changes from liquid to gas and vice versa?



What are the names for the state changes from liquid to gas and vice versa?

Liquid → gas: Evaporation

Gas → liquid: Condensation



Describe the forces between particles in
solids, liquids and gases



Describe the forces between particles in solids, liquids and gases

Solids: Strong forces of attraction between particles which keep them in fixed positions.

Liquids: Weaker attractive forces than in solids.

Gases: Weakest intermolecular forces so particles move randomly.



How does a physical change differ from a chemical change?



How does a physical change differ from a chemical change?

A physical change involves changes in the forces between particles. The particles themselves remain the same and the chemical properties remain the same.

A chemical change is different as it affects the chemical properties of the substance.



True or false?
'Physical changes are relatively easy to reverse'



True or false?

‘Physical changes are relatively easy to reverse’

TRUE

Relatively easy to reverse since no new product is formed during the changes of state.



Describe what happens, in terms of particles, when a solid is heated and melts into a liquid



Describe what happens, in terms of particles, when a solid is heated and melts into a liquid

When heated, particles absorb thermal energy and convert it into kinetic energy. The particles in the solid vibrate more, causing the solid to expand until the structure breaks and the solid turns into a liquid.



Describe what happens, in terms of particles, when a liquid is heated and evaporates into a gas



Describe what happens, in terms of particles, when a liquid is heated and evaporates into a gas

When heated, the particles in a liquid expand. Some particles on the surface gain sufficient energy to overcome the intermolecular forces and evaporate. At the boiling point, all of the liquid particles gain enough energy to evaporate.



Substance A melts at -200°C and boils at -183°C . What state is A at -174°C ?



Substance A melts at -200°C and boils at -183°C .
What state is A at -174°C ?

Liquid



Substance B melts at -5°C and boils at 23°C . What state is A at -7°C ?



Substance B melts at -5°C and boils at 23°C . What state is A at -7°C ?

Solid



Why do solids, liquids and gases expand when heated?



Why do solids, liquids and gases expand when heated?

When a substance is heated the molecules vibrate faster. This causes the space between the atoms to increase.



What is an advantage of the current particle model?



What is an advantage of the current particle model?

It provides a simple understandable model to explain the three states of matter.



‘Particles in the particle model are represented by _____’

Fill in the gaps
(Higher only)



'Particles in the particle model are represented by

,

Fill in the gaps (Higher only)

Inelastic spheres



What are the limitations of the particle
model?
(Higher only)



What are the limitations of the particle model?

(Higher only)

- Doesn't take into account forces of attraction between particles. The amount of energy required to cause a change of state depends on these forces.
- Doesn't take into account the size of particles and space between them. The nature of particles depends on the structure and bonding of a substance.



Liquid A has a higher boiling point than liquid B. What does this tell you about the forces between the particles in liquid A?

(Higher only)



Liquid A has a higher boiling point than liquid B.
What does this tell you about the forces between the particles in liquid A? (Higher only)

Liquid A has greater forces of attraction between the particles.



What is meant by the terms element and compound?



What is meant by the terms element and compound?

Elements are substances made up of only one type of atom.

Compounds are made up of atoms of different elements.



What are the three subatomic particles in an atom?



What are the three subatomic particles in an atom?

Protons

Neutrons

Electrons



Who described atoms as ‘solid spheres’?



Who described atoms as 'solid spheres'?

John Dalton



What was Dalton's theory?



What was Dalton's theory?

- Atoms cannot be created, divided or destroyed.
- Atoms of the same element are exactly the same and atoms of different elements are different.
- Atoms join with other atoms to make new substances.



What discovery caused the Dalton model of an atom to change?



What discovery caused the Dalton model of an atom to change?

The discovery of subatomic particles



How did JJ Thomson discover the electron?



How did JJ Thomson discover the electron?

Thomson conducted an experiment using a cathode ray tube. The beam moved towards the positively charged plate so he knew the particles must have a negative charge.



Describe the atomic model proposed by
JJ Thomson



Describe the atomic model proposed by JJ Thomson

The plum pudding model:

Negative electrons scattered through a positively charged material.



Who designed and carried out the gold foil experiment?



Who designed and carried out the gold foil experiment?

Ernest Rutherford designed the experiment.

Geiger and Marsden carried out the experiment.



What did Rutherford discover from his
gold foil experiment?



What did Rutherford discover from his gold foil experiment?

He shot a beam of positively charged particles into a sheet of gold foil.

- Most particles passed straight through indicating that atoms were mostly empty space.
- A few particles were deflected and a few bounced directly back showing that there must be a positively charged nucleus.



Describe Rutherford's model of the atom



Describe Rutherford's model of the atom

- Mass is concentrated in the central nucleus.
- Mostly empty space.
- Electrons travel in random paths around the nucleus.



What change did Niels Bohr propose to the nuclear model of an atom?



What change did Niels Bohr propose to the nuclear model of an atom?

Electron shells around the nucleus.



Describe the structure of an atom



Describe the structure of an atom

- A small central nucleus made up of protons and neutrons (positively charged).
- Electrons orbit (move around) the nucleus in shells (negatively charged).



Where is the mass of the atom concentrated?



Where is the mass of the atom concentrated?

In the nucleus



Compare the sizes of the nuclear radius
and the atomic radius



Compare the sizes of the nuclear radius and the atomic radius

The nuclear radius is much smaller than the atomic radius.



Compare the typical size of atoms and small molecules



Compare the typical size of atoms and small molecules

Atoms and small molecules are both incredibly small.

Small molecules are larger than atoms because they are made of atoms.

The typical atomic radii and bond length are in the order of 10^{-10}m .



What are the relative masses of a proton, neutron and electron?



What are the relative masses of a proton, neutron and electron?

Proton: 1

Neutron: 1

Electron: $1/1836$



What are the relative charges of a proton, neutron and electron?



What are the relative charges of a proton, neutron and electron?

Proton: +1

Neutron: 0

Electron: -1



What is an ion?



What is an ion?

An atom or molecule with a positive or negative charge.



How is an ion formed?



How is an ion formed?

When an atom or molecule gains or loses electron(s).

Positive ions are formed when an electron is lost.

Negative ions are formed when an electron is gained.



What does the atomic number tell you about an element?



What does the atomic number tell you about an element?

The atomic number is unique to each element and tells you the number of protons an element has.



What is the mass number?



What is the mass number?

The combined total of the number of protons and neutrons in the nucleus of an atom of an element.



What are isotopes?



What are isotopes?

Atoms of the same element with the same number of protons but a different number of neutrons.



How does the atomic number and mass number differ between isotopes of the same element?



How does the atomic number and mass number differ between isotopes of the same element?

Atomic number is the same as an element always has the same number of protons.

Mass number is different as there are different numbers of neutrons.



Why do atoms contain equal numbers of protons and electrons?



Why do atoms contain equal numbers of protons and electrons?

Atoms have a stable overall charge of 0.

Protons are positively charged and electrons are negatively charged so they must be present in equal numbers for charges to balance.



How can you calculate the number of neutrons, given the mass number and atomic number of an element?



How can you calculate the number of neutrons, given the mass number and atomic number of an element?

Number of neutrons =
mass number - atomic number



Boron has the atomic number 5 and mass number 11. How many protons, electrons and neutrons does boron have?



Boron has the atomic number 5 and mass number 11. How many protons, electrons and neutrons does boron have?

5 protons

5 electrons

6 neutrons



Sodium has the atomic number 11 and mass number 23. How many protons, electrons and neutrons does the Na^+ ion have?



Sodium has the atomic number 11 and mass number 23. How many protons, electrons and neutrons does the Na^+ ion have?

11 protons

10 electrons (one has been to form the positive ion)

12 neutrons

